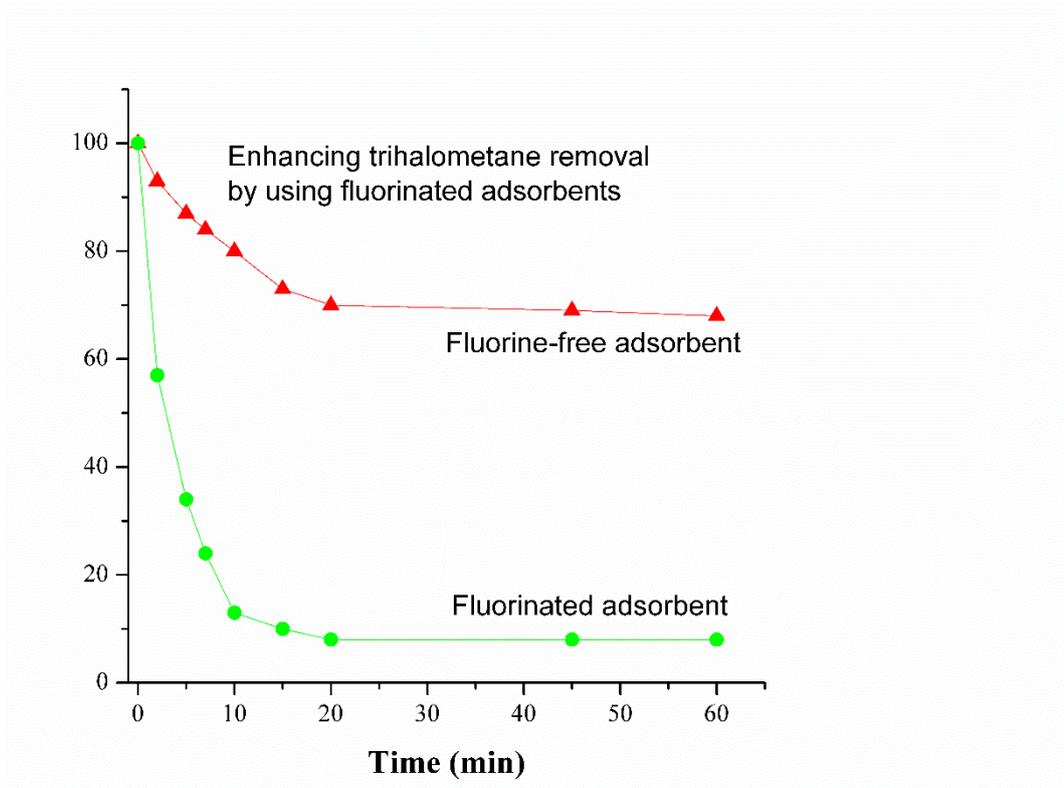


# Graphical abstract



CAPTIONS TO FIGURES

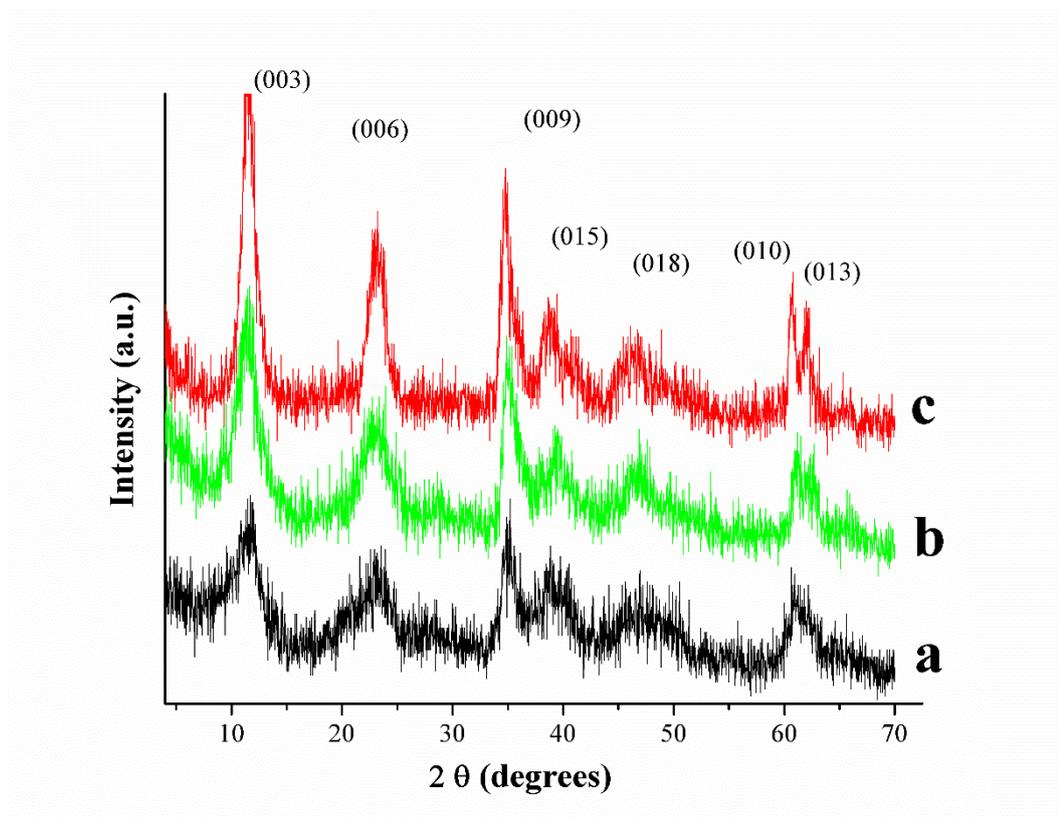


Figure 1. X-ray diffraction patterns of LDH as synthesized: (a) HT, (b) HT-10F and (c) HT-25F. Peak labels indicate Miller index of JPCDS card 22-0700.

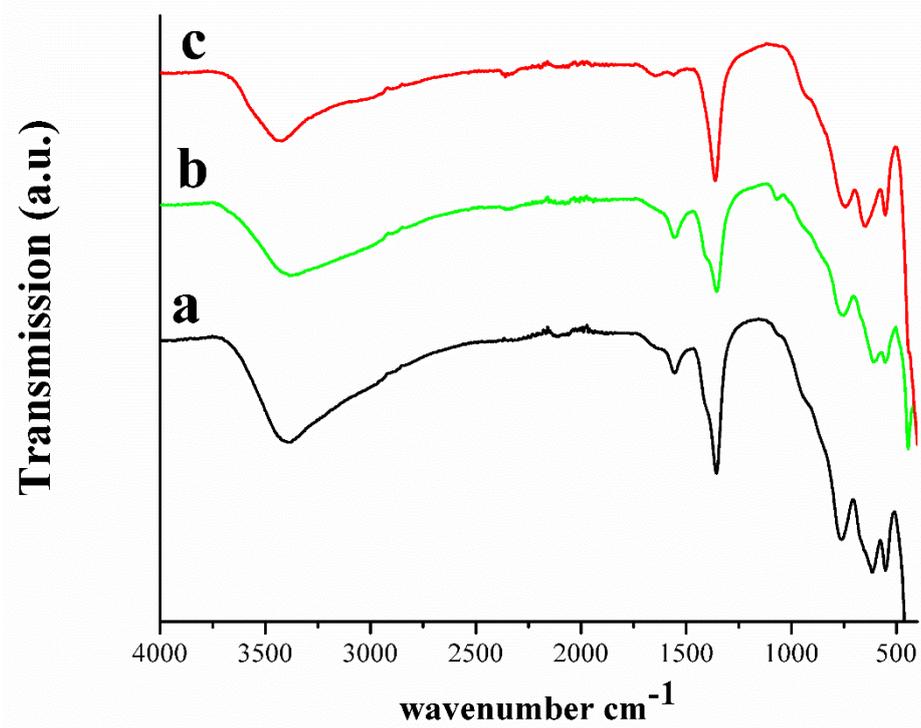


Figure 2. Infrared spectra of LDH as synthesized: (a) HT, (b) HT-10F and (c) HT-25F.

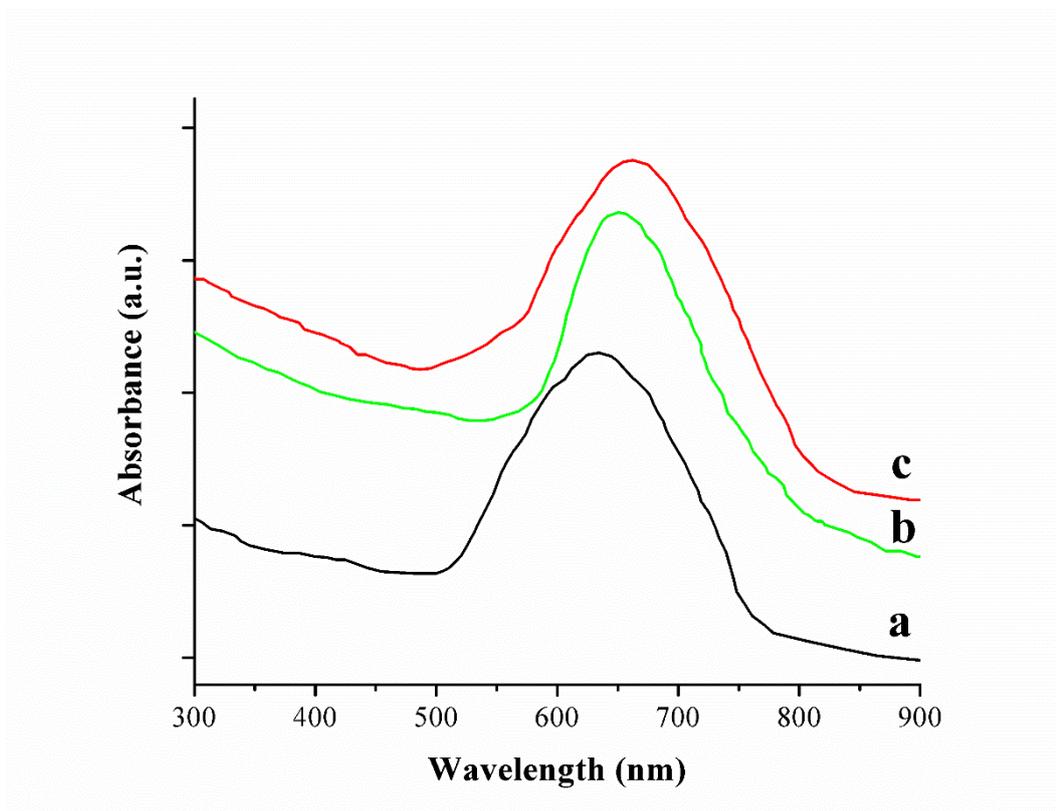


Figure 3. UV-vis spectra of the solvatochromic dye 4-tert-butyl-2-(dicyanomethylene)-5-[4-(diethylamino)benzylidene]- $\Delta^3$ -thiazoline adsorbed onto (a) HT, (b) HT-10F and (c) HT-25F.

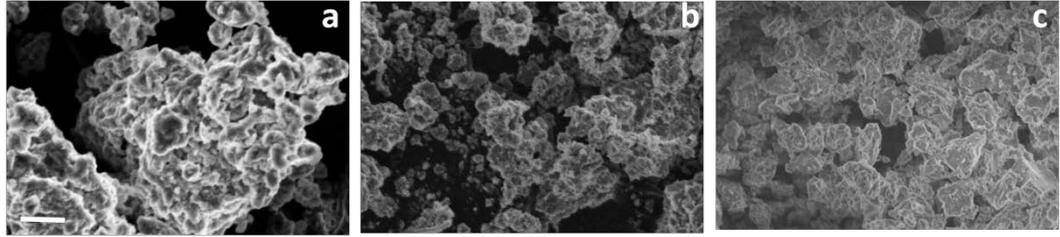


Figure 4. SEM image of (a) HT, (b) HT-10F and (c) HT-25F. The bar scale in image (a) corresponds to 20  $\mu\text{m}$  and it is valid for all three images.

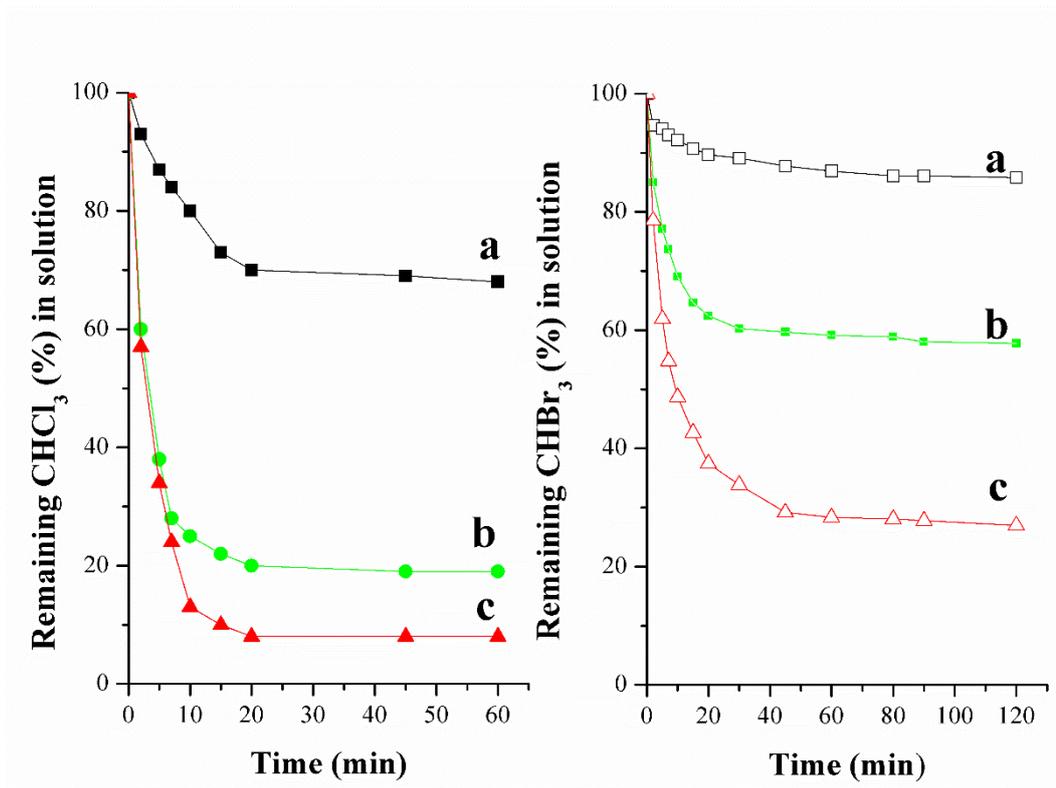


Figure 5. Concentration progress of CHCl<sub>3</sub> (left) and CHBr<sub>3</sub> (right) after contact with different adsorbents (a) HTC, (b) HT-10FC and (c) HT-25FC. Initial concentration of CHCl<sub>3</sub> and CHBr<sub>3</sub> was  $5 \times 10^{-3}$  and  $4 \times 10^{-4}$  M, respectively. Adding of C to code sample means that adsorbent was thermal treated at 350 °C for 8 h. The ratio mass of adsorbent to volume of solution was fixed to 12.5 mg/ml.

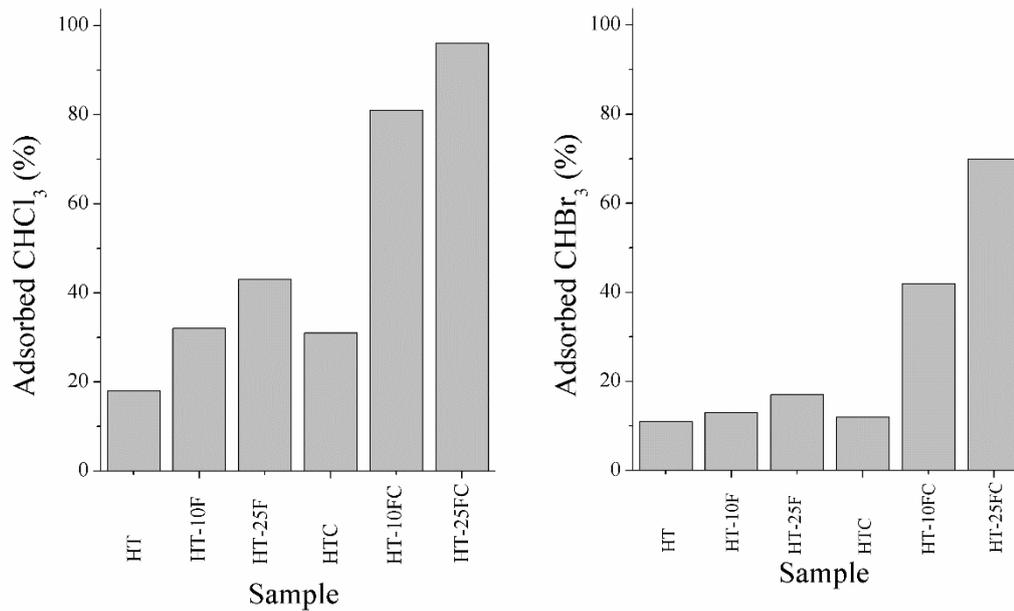


Figure 6.  $\text{CHCl}_3$  (left) and  $\text{CHBr}_3$  (right) removal efficiency through adsorption on different adsorbents. The equilibrium time was taken as 30 and 90 minutes for  $\text{CHCl}_3$  and  $\text{CHBr}_3$ , respectively.

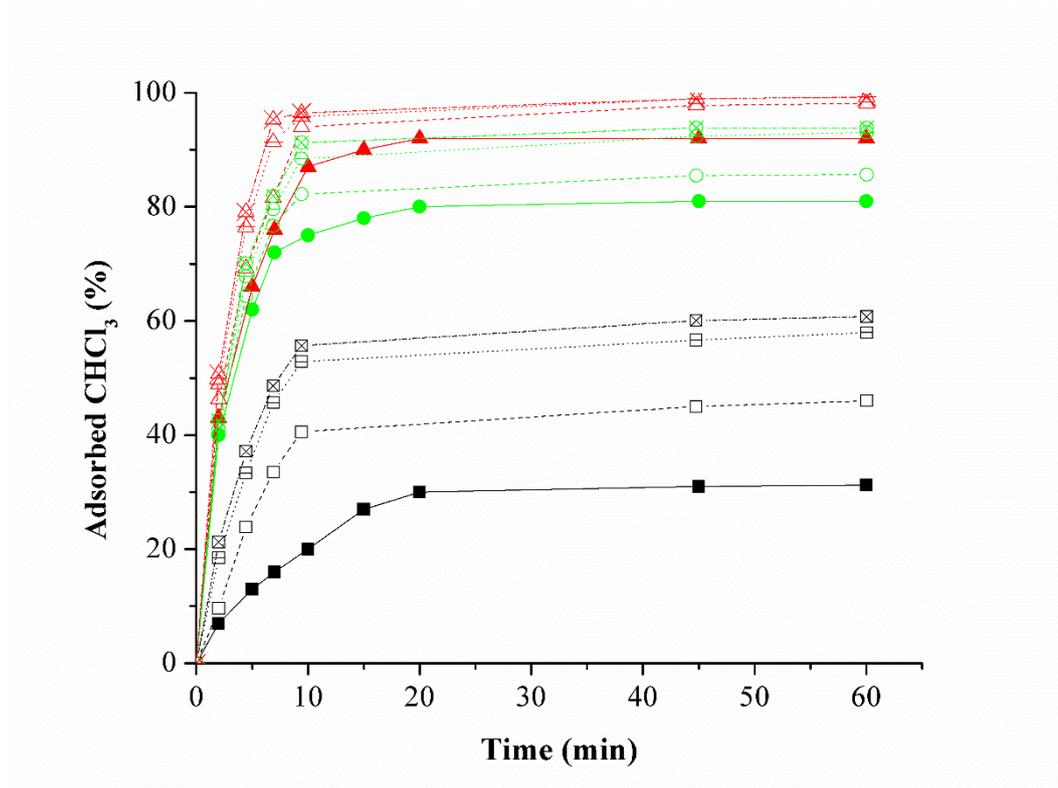


Figure 7. CHCl<sub>3</sub> adsorption onto different adsorbents (squares) HTC, (circles) HT-10FC and (triangles) HT-25FC, starting with different initial CHCl<sub>3</sub> concentration, 5 × 10<sup>-3</sup> M (solid lines), 2 × 10<sup>-2</sup> (dashed lines), 3 × 10<sup>-2</sup> (dotted lines), 5 × 10<sup>-2</sup> (dotted-dashed lines).

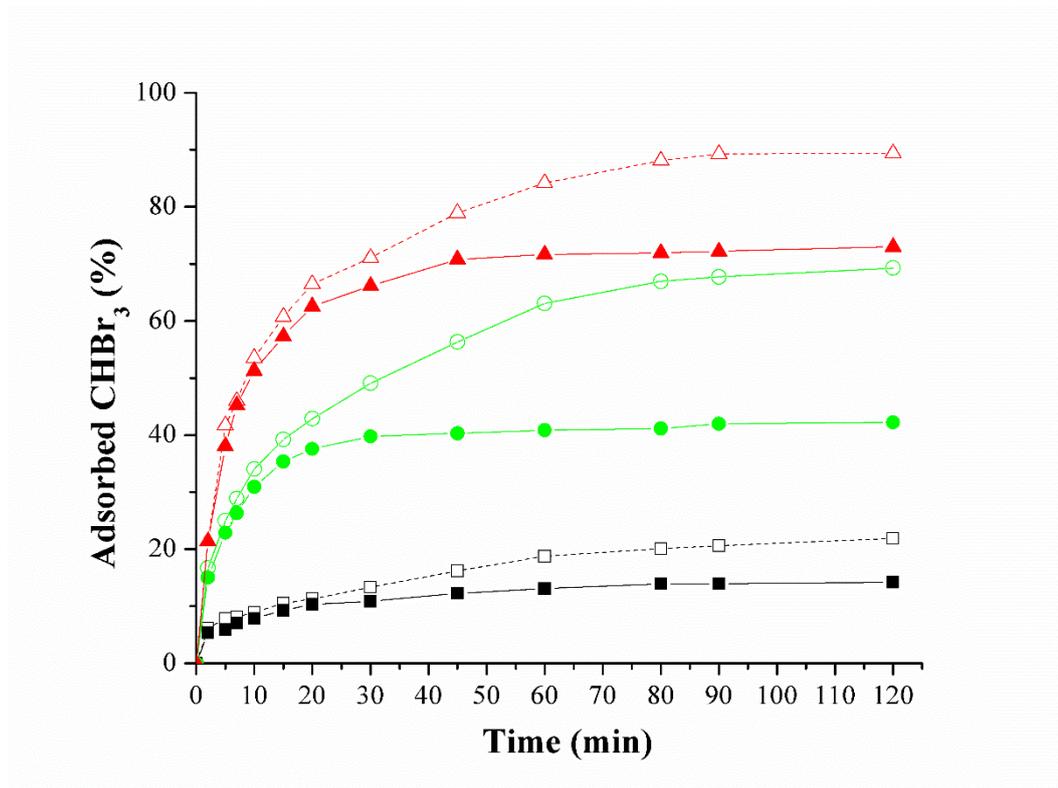


Figure 8. CHBr<sub>3</sub> adsorption onto different adsorbents (squares) HTC, (circles) HT-10FC and (triangles) HT-25FC, starting with different initial CHCl<sub>3</sub> concentration,  $4 \times 10^{-4}$  M (solid lines),  $5 \times 10^{-3}$  (dashed lines).

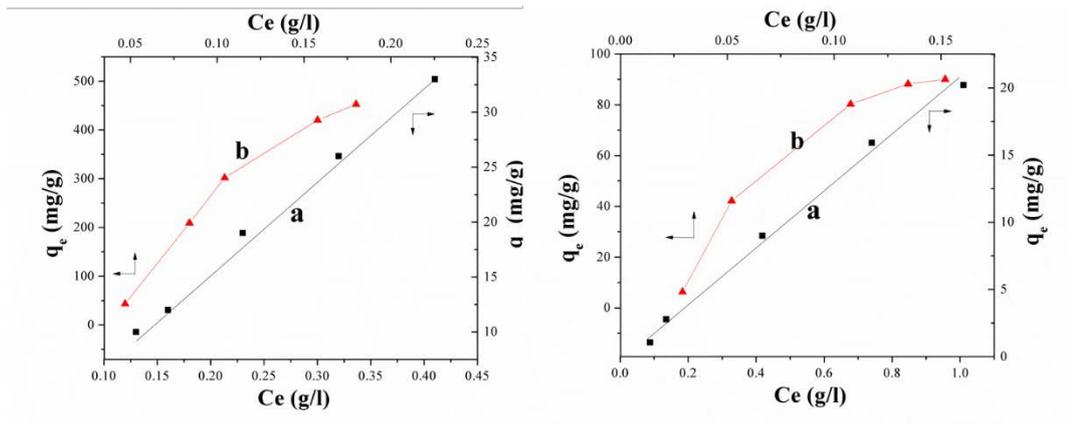


Figure 9. Adsorption isotherms of  $CHCl_3$  (left) and  $CHBr_3$  (right) onto (a) HTC, and (b) HT-25FC.  $q_e$  represents the amount of THM per unit mass of adsorbent and  $C_e$  the equilibrium concentration of the remaining THM in solution. The equilibrium time was taken as 30 and 90 minutes for  $CHCl_3$  and  $CHBr_3$ , respectively.